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APPLICATION NO.	FILING DATE	FIRST NAMED INVENTOR	ATTORNEY DOCKET NO.	CONFIRMATION NO.
10/776,063	02/11/2004	John S. Benco	1657/2052	2112
29932	7590	12/07/2005	EXAMINER	
<b>PAULA CAMPBELL EVANS</b> SONNENSCHEIN NATH & ROSENTHAL LLP ONE INTERNATIONAL PLACE BOSTON, MA 02110				SODERQUIST, ARLEN
		ART UNIT		PAPER NUMBER
		1743		

DATE MAILED: 12/07/2005

Please find below and/or attached an Office communication concerning this application or proceeding.

<b>Office Action Summary</b>	Application No.	Applicant(s)
	10/776,063	BENCO ET AL.
	Examiner	Art Unit
	Arlen Soderquist	1743

-- The MAILING DATE of this communication appears on the cover sheet with the correspondence address --  
**Period for Reply**

A SHORTENED STATUTORY PERIOD FOR REPLY IS SET TO EXPIRE 3 MONTH(S) FROM THE MAILING DATE OF THIS COMMUNICATION.

- Extensions of time may be available under the provisions of 37 CFR 1.136(a). In no event, however, may a reply be timely filed after SIX (6) MONTHS from the mailing date of this communication.
- If the period for reply specified above is less than thirty (30) days, a reply within the statutory minimum of thirty (30) days will be considered timely.
- If NO period for reply is specified above, the maximum statutory period will apply and will expire SIX (6) MONTHS from the mailing date of this communication.
- Failure to reply within the set or extended period for reply will, by statute, cause the application to become ABANDONED (35 U.S.C. § 133). Any reply received by the Office later than three months after the mailing date of this communication, even if timely filed, may reduce any earned patent term adjustment. See 37 CFR 1.704(b).

**Status**

- 1) Responsive to communication(s) filed on 20 September 2005.
- 2a) This action is **FINAL**.                                    2b) This action is non-final.
- 3) Since this application is in condition for allowance except for formal matters, prosecution as to the merits is closed in accordance with the practice under *Ex parte Quayle*, 1935 C.D. 11, 453 O.G. 213.

**Disposition of Claims**

- 4) Claim(s) 1-53 is/are pending in the application.
- 4a) Of the above claim(s) \_\_\_\_\_ is/are withdrawn from consideration.
- 5) Claim(s) \_\_\_\_\_ is/are allowed.
- 6) Claim(s) 1-53 is/are rejected.
- 7) Claim(s) \_\_\_\_\_ is/are objected to.
- 8) Claim(s) \_\_\_\_\_ are subject to restriction and/or election requirement.

**Application Papers**

- 9) The specification is objected to by the Examiner.
- 10) The drawing(s) filed on 11 February 2004 is/are: a) accepted or b) objected to by the Examiner.  
 Applicant may not request that any objection to the drawing(s) be held in abeyance. See 37 CFR 1.85(a).  
 Replacement drawing sheet(s) including the correction is required if the drawing(s) is objected to. See 37 CFR 1.121(d).
- 11) The oath or declaration is objected to by the Examiner. Note the attached Office Action or form PTO-152.

**Priority under 35 U.S.C. § 119**

- 12) Acknowledgment is made of a claim for foreign priority under 35 U.S.C. § 119(a)-(d) or (f).  
 a) All    b) Some \* c) None of:  
 1. Certified copies of the priority documents have been received.  
 2. Certified copies of the priority documents have been received in Application No. \_\_\_\_\_.  
 3. Copies of the certified copies of the priority documents have been received in this National Stage application from the International Bureau (PCT Rule 17.2(a)).

\* See the attached detailed Office action for a list of the certified copies not received.

**Attachment(s)**

- 1) Notice of References Cited (PTO-892)
- 2) Notice of Draftsperson's Patent Drawing Review (PTO-948)
- 3) Information Disclosure Statement(s) (PTO-1449 or PTO/SB/08)  
 Paper No(s)/Mail Date \_\_\_\_\_.
- 4) Interview Summary (PTO-413)  
 Paper No(s)/Mail Date. \_\_\_\_\_.
- 5) Notice of Informal Patent Application (PTO-152)
- 6) Other: \_\_\_\_\_.

1. The following is a quotation of the first paragraph of 35 U.S.C. 112:

The specification shall contain a written description of the invention, and of the manner and process of making and using it, in such full, clear, concise, and exact terms as to enable any person skilled in the art to which it pertains, or with which it is most nearly connected, to make and use the same and shall set forth the best mode contemplated by the inventor of carrying out his invention.

2. Claims 1-53 are rejected under 35 U.S.C. 112, first paragraph, because the specification, while being enabling for the specifically disclosed compound, does not reasonably provide enablement for the variation in the R, X, Y and Z groups as set forth in the independent claims. The specification does not enable any person skilled in the art to which it pertains, or with which it is most nearly connected, to make and use the invention commensurate in scope with these claims. First in the response filed April 4, 2005, applicant has argued that “devising a compound that will exclusively bind to one of these ions is an *unpredictable art*” (emphasis added, see the last paragraph of page 9 in the response). The above statement is after applicant has characterized the claimed compound as having the components carefully selected to optimize lithium detection from sodium and potassium ions, which are characterized by applicant as being “extremely similar in structure and properties”. Page 6, lines 17-22 of the instant specification describe some properties of calixarenes and state that by “functionally modifying either the upper and/or lower rims it is possible to prepare various derivatives with differing selectivities for various guest ions and small molecules”. Thus the selectivity of the calixarene is dependent on the way in which the upper and or lower rims are functionalized. Each if the R, X, Y and Z groups are a functionalization of either the upper or lower rims of the claimed calixarene and contain a plurality of different elements that span a wide variety of properties. Thus due to the unpredictability that applicant has argued, the single example given in the instant specification fails to provide enablement for the large variety of structures and properties claimed in the R, X, Y, and Z groups.

3. The following is a quotation of 35 U.S.C. 103(a) which forms the basis for all obviousness rejections set forth in this Office action:

(a) A patent may not be obtained though the invention is not identically disclosed or described as set forth in section 102 of this title, if the differences between the subject matter sought to be patented and the prior art are such that the subject matter as a whole would have been obvious at the time the invention was made to a person having ordinary skill in the art to which said subject matter pertains. Patentability shall not be negated by the manner in which the invention was made.

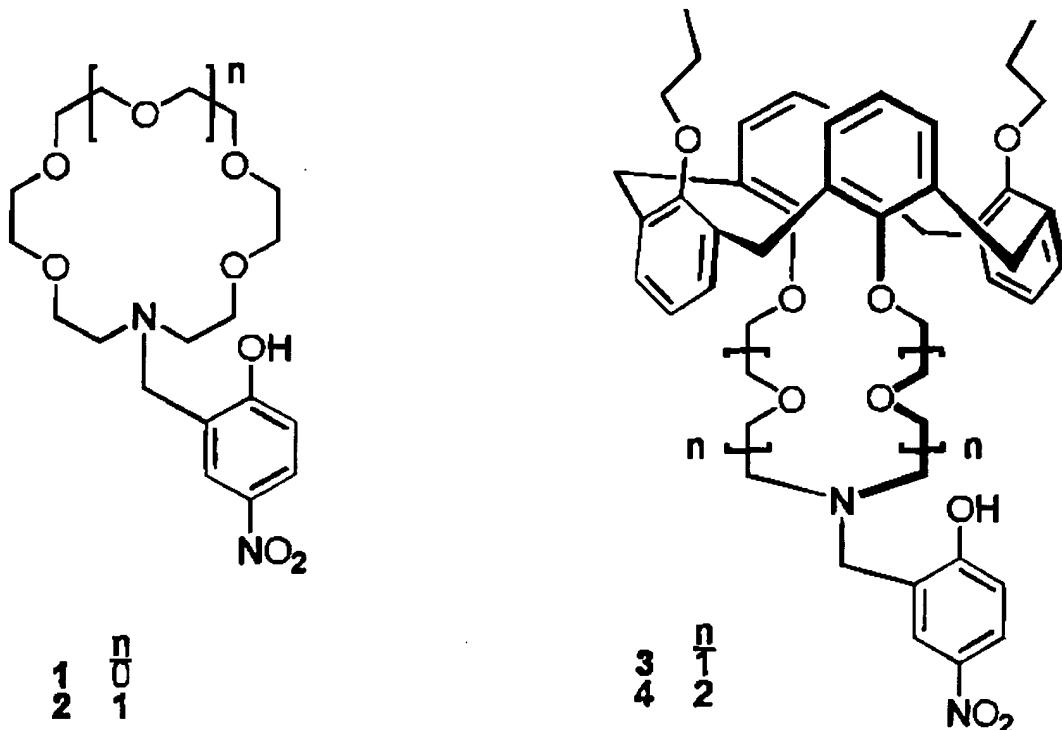
The factual inquiries set forth in *Graham v. John Deere Co.*, 383 U.S. 1, 148 USPQ 459 (1966), that are applied for establishing a background for determining obviousness under 35 U.S.C. 103(a) are summarized as follows:

1. Determining the scope and contents of the prior art.
  2. Ascertaining the differences between the prior art and the claims at issue.
  3. Resolving the level of ordinary skill in the pertinent art.
  4. Considering objective evidence present in the application indicating obviousness or nonobviousness.
4. Claims 1-53 are rejected under 35 U.S.C. 103(a) as being unpatentable over Benco (*Journal of Photochemistry and Photobiology A*, 2002) or Kim (J. Org Chem.) in view of Pacey (US 4,659,815) and Barnard (US 6,417,005 or WO 97/39337).

In the paper Benco teaches 9-anthryl-substituted azacrown ether covalently linked to a 1,3-alternate calix[4]arene as a fluoroionophore for the detection of potassium ions. *N*-(9-methyl-anthracene)-25,27-bis(1-propyloxy)calix[4]arene azacrown-5 (**II**) and its model compound *N*-(9-anthrylmethyl)aza-18-crown-6 (**I**) were synthesized and tested as fluoroionophores for the selective detection of potassium ions with a view to the use of **II** in the fabrication of potassium ion sensors. Compound **II** consists of a 1,3-alternate calix[4]arene group covalently linked to an azacrown ether that is *N*-substituted with a fluorescent anthracene group. This compound acts as an 'off-on' fluorescent indicator for ion complexation. In dichloromethane solution, compound **II** exhibits good sensitivity to potassium ions and forms a 1:1 fluoroionophore-ion complex. Studies demonstrate that **II** is selective for potassium over other alkali metal cations, with excellent selectivity over sodium and lithium ( $\log K_{K,Na} \sim \log K_{K,Li} \leq 3.5$ ) and moderate selectivity over rubidium and cesium ( $\log K_{K,Rb} \sim \log K_{K,Cs} \sim 1$ ). Sensitivity of **II** to potassium is considerably enhanced in dichloromethane in comparison to methanol/dichloromethane mixtures, presumably due to two effects: a hydrogen-bonding interaction of methanol with the azacrown nitrogen atom, and poor solvation of the ion by dichloromethane, the latter creating a driving force for complexation. On page 37 there is a discussion of selectivity. The first sentence of this section teaches that the eventual intention is to use molecules similar to **II** as a sensor. The next sentence teaches that given the structural similarities between **II** and 1,3 alternate calix[4]arenas, it is reasonable to expect similar binding properties. The next sentence teaches that they expect metal complexation in **II** to be governed

by electrostatic interactions (primarily cation- $\pi$  interactions). That sentence also teaches that the selectivity is primarily controlled by a size fit effect and steric effects from the propyl substituents appended to the two rotated aryl rings of the calix[4]arene. Page 33 discusses the use of similar compounds for other metal cations pointing to this paper directed generally at sensing cations through complexing structures. In this respect Benco is directed to identification of a species of the genus. In the first full paragraph of page 34 Benco teaches that they synthesized a compound of structure II (and its model compound of structure I) because *it combines the known optical response of the anthryl azacrown 6 to potassium ions with expected enhanced selectivity provided by coupling the azacrown to a 1,3-alternate calyx[4]arene*. Benco does not teach an azacrown calix[4]arene sized to capture lithium or structures related to sensor formation.

In the paper Kim teaches the synthesis and metal ion complexation studies of proton-ionizable calix[4]azacrown ethers in the 1,3-alternate conformation. A series of novel N-chromogenic calix[4]arene azacrown ethers shown below as compounds 3-4 were synthesized. The azacrown compounds 1-2 were also synthesized and the metal complexation/extraction behavior was measured. Page 2386 discusses the known use of both of these types of structures for determination of metal ions. Structure 1 was found to have lithium selectivity while



compound 2 showed potassium selectivity due to size agreement between the metal ion and the cavity of the corresponding azacrown ether. The paragraph bridging pages 2386-2387 discusses calixcrown ethers as 3-D molecular building blocks for receptors of metal cations. For cesium in particular, a calixarene crown derivative is exceptionally good due to the complexation of cesium ion not only with the crown ether but also with the two aromatic rings (cation/π-interaction) when fixed in the 1,3-alternate conformation. The first full paragraph of page 2387 teaches that based in this, **it is possible that the combination of *N*-chromogenic azacrown ether and calixcrown ether would result in an optimized structure for metal ion encapsulation** due to (1) electrostatic interactions between the metal ion and both the oxygens and a nitrogen as electron donors, (2) π-metal interactions between the metal ion and two rotated aromatic nuclei of the 1,3-alternate calixarene, and (3) an extra pendant chromogenic group attached to nitrogen, which can promote metal complexation by 3-D encapsulation under basic conditions. With this in mind, they sought to synthesize chromogenic calixazacrown ethers and to investigate their complexation behavior **toward alkali metal ions** through bulk liquid membrane, solvent extraction, and  $^1\text{H}$  NMR studies. The chromogenic calix[4]arene azacrown ethers were prepared

in moderate yields by reacting dipropoxybis(chlorooxapentyl)calix[4]arenes with p-toluenesulfonamide in the presence of potassium carbonate; reductive removal of the tosyl group and alkylation of the amine with 2-hydroxy-5-nitrobenzyl bromide. Tables 1-3 show the results of the tests with high transport selectivity for potassium over other metal ions for compound 3 as shown by two-phase extraction, bulk liquid membrane, and <sup>1</sup>H NMR studies on a ligand-metal complex. It is assumed that the OH of the chromogenic group attached on nitrogen can assist the complexation by encapsulation of the metal. Compound 3-4 are based on the structure of compounds 1-2. Page 2386 and the first full paragraph of page 2387 are directed toward metal ion encapsulation and detection. While the specific compounds were selective toward a species of the genus, the reference is not so limited as shown by the first full paragraph of page 2387 as noted above. Kim does not teach an azacrown calix[4]arene sized to capture lithium or structures related to sensor formation.

In the patent Pacey teaches chromogenic aza-12-crown-4 ethers used for the spectrophotometric determination of lithium ion in aqueous solutions. The compounds are particularly useful for the analysis of lithium in the presence of Na<sup>+</sup>, a situation common in biological and geological systems. The compounds [e.g., 1-(2-oxy-5-nitrobenzyl)-1-hydro-1-aza-4,7,10-trioxacyclododecane] are similar in structure to compound I of Benco except they have a smaller crown and a colorimetric group appended to the crown. Column 1, lines 30-43 teach that selective reagents permitting the isolation of a particular ion from a complex matrix or mixture of ions are of interest to those in the chemical or bio-chemical analytical fields. When the matrix in question includes two or more cations of the Group I metals, it is often difficult to selectively isolate one of the Group I cations from the mixture without interference from other cations in the system. With respect to biological systems, such as blood serum, which contains a relatively large amount of sodium ion, a reagent having the ability to selectively (and quantitatively) extract lithium ion in the presence of sodium ion is of interest for bio-chemical assays. Some examples of prior art molecules are taught with drawbacks being explained. Column 3, lines 11-19 and example 9 teach the selectivity of the molecules of the invention when lithium is complexed in the presence of a large excess of sodium. Example 7 shows the ability of the reagents to complex lithium in the presence of a potassium hydroxide buffer and

example 11 shows the ability to measure lithium in blood and urine samples. These results shown in table 2 show excellent correlation with the known amount of lithium present.

In the patent and published application Barnard teaches covalently immobilized fluoroionophores as optical metal ion sensors. Since both are members of the same patent family only the US patent will be described but corresponding disclosure also exists in the published application. Fluoroionophores that are fictionalized with reactive groups may be covalently bound to inorganic or organic carrier materials and are used as active components in polymer membranes of optical sensors for the detection of ions. The sensors are distinguished by a long usable life and a high degree of sensitivity. In the formula taught, and ionophore, I, can have a number of forms and are natural or synthetic organic compounds that contain a plurality of mostly alternating electron-rich hetero atoms such as, for example, S, N and especially O, in an open-chained or cyclic carbon chain and that enable the ions to be measured to be selectively complexed. Column 2, line 60 to column 3, line 7 teach these ionophores derived from substances that have an oligoether, polyether, oligoester, polyester, oligoamide or polyamide structure. Examples of such suitable substances may be crown ethers, coronandenes, cryptandenes, **calixarenes**, podandene or derivatives thereof, also cyclic peptides; for example valinomycin, nonactin, peptides such as gramicidin, and peptides which in the presence of the ion to be determined change their secondary, tertiary or quaternary structure for bonding the ion; it is also possible to use tetrahydrofuran-containing macrolides bonded via ester bridges, and analogous substances that are able to regulate transport in biological systems, or cyclic oligosaccharides, such as, for example, cyclodextrins or cyclophanes. Columns 5-8 show the structure of several of these with structure (IV) being a calix[4]arene. Columns 8-11 show various examples of the fluorophores used in the fluoroionophores. Columns 13-18 specifically show structures of three calix[4]arene fluoroionophores. Columns 20-22 teach a wide variety of carrier materials. Column 29 teaches that the compositions taught may be applied to suitable support materials. The support is preferably transparent and may be formed, for example, from a plastics material, such as, for example, polycarbonate or acrylic glass, mineral materials, metal oxides or glass, and may be of any shape, for example in the form of plates, cylinders, tubes, strips or fibers. The optical range in which the material as sensor can be excited extends from the ultraviolet range to the infrared range. The immobilized fluorophore-ionophores have very

suitable absorption and emission wavelength ranges that allow the use of known economically priced low-energy light sources, for example halogen or xenon lamps or light-emitting diodes. Commercially obtainable optical fibers may be used in the excitation and detection. Column 30 teaches that one very important advantage is that the immobilized fluoroionophores can carry out the analyses being substantially independent of pH and therefore may be carried out directly in body fluids such as blood, urine and serum. Among the many cations that the compounds can be used for, the alkali metal ions including lithium are preferred cations (column 30).

It would have been obvious to one of ordinary skill in the art at the time the invention was made to modify the azacrown portion of the Benco or Kim fluoroionophore to correspond to that of the difference between the model azacrown of Benco or Kim and the similar azacrown of Pacey because of the recognition by Benco or Kim that the binding is related to the size fit and other types of effects and the teaching in Pacey that lithium is an important analyte that can be complexed by the smaller azacrown with a high selectivity with respect to the complexation of sodium and other alkali metals in a biological sample. It would have been obvious to one of ordinary skill in the art at the time the invention was made to incorporate the fluoroionophore of Benco into a carrier material and or onto a support as taught by Barnard because of the advantages in carrying out the analysis as taught by Barnard.

5. Claims 1-53 are rejected under 35 U.S.C. 103(a) as being unpatentable over Pacey in view of Benco or Kim and Barnard as described above. Pacey does not teach the azacrown attached to a calix[4]arene or structures related to sensor formation.

It would have been obvious to one of ordinary skill in the art at the time the invention was made to modify the azacrown of Pacey with a calyx[4]arene as taught for the Benco or Kim fluoroionophore to correspond to that of the difference between the model azacrown of Benco or Kim and the similar azacrown of Pacey because of the recognition by Benco or Kim that the binding is related to the size fit and other types of effects, the combination of an azacrown with a calix[4]arene has advantages relative to selectivity optimization of the metal ion encapsulation process as taught by Benco or Kim and the teaching in Pacey that lithium is an important analyte that can be complexed by the smaller azacrown with a high selectivity with respect to the complexation of sodium and other alkali metals in a biological sample. It would have been obvious to one of ordinary skill in the art at the time the invention was made to incorporate the

fluoroionophore of Pacey into a carrier material and or onto a support as taught by Barnard because of the advantages in carrying out the analysis as taught by Barnard.

6. The nonstatutory double patenting rejection is based on a judicially created doctrine grounded in public policy (a policy reflected in the statute) so as to prevent the unjustified or improper timewise extension of the "right to exclude" granted by a patent and to prevent possible harassment by multiple assignees. See *In re Goodman*, 11 F.3d 1046, 29 USPQ2d 2010 (Fed. Cir. 1993); *In re Longi*, 759 F.2d 887, 225 USPQ 645 (Fed. Cir. 1985); *In re Van Ornum*, 686 F.2d 937, 214 USPQ 761 (CCPA 1982); *In re Vogel*, 422 F.2d 438, 164 USPQ 619 (CCPA 1970); and, *In re Thorington*, 418 F.2d 528, 163 USPQ 644 (CCPA 1969).

A timely filed terminal disclaimer in compliance with 37 CFR 1.321(c) may be used to overcome an actual or provisional rejection based on a nonstatutory double patenting ground provided the conflicting application or patent is shown to be commonly owned with this application. See 37 CFR 1.130(b).

Effective January 1, 1994, a registered attorney or agent of record may sign a terminal disclaimer. A terminal disclaimer signed by the assignee must fully comply with 37 CFR 3.73(b).

7. Claims 1-53 are rejected under the judicially created doctrine of obviousness-type double patenting as being unpatentable over claims 1-28 of U.S. Patent No. 6,660,526 in view of Benco, Pacey and Barnard as explained above. The patented claims are directed to a molecule, device and method that encompass the molecule, device and method taught by Benco above, therefore they differ from the instant claims as the Benco reference differs from the instant claims. Thus for the reasons given above, the instant claims are obvious in view of the patented claims.

8. Applicant's arguments filed September 20, 2005 have been fully considered but they are not persuasive. First applicant's arguments in the April 4, 2005 response relative to the unpredictability of designing a compound to detect lithium, sodium or potassium raised the question of the scope of enablement since the instant application shows a single compound, yet the scope of the claims appears to cover a large number of compounds with the respective R, X, Y, and Z groups covering a wide variety of variety of moieties with significantly varying properties. For example, just within the R group a hydrogen, an alkyl amine, a carboxylic acid, a pyridinyl group, a triazinyl group and a benzothiol group have different properties and are capable of appearing at all of the positions of the molecule so designated. Applicant appears to be arguing that either all will have selectivity of lithium over sodium or potassium or that only the combinations from the exceptionally large group of possibilities that have the claimed lithium selectivity property are within the scope of the claims. With the differences in properties

of each of the above R groups, examiner fails to see where the expectation that compounds having each of the different R groups at all of the R group locations would be expected to have a selectivity to lithium when compared to sodium and potassium. This coupled with the statement that the functionalization of the upper and/or lower rims of the calixarene molecule can change the selectivity for various guest ions and small molecules found in the instant specification clearly raises a question of enablement of the full scope of the claims. It also raises a question regarding the ability of synthesizing all of the myriad possibilities including having R be different members of the claimed grouping at different sites of the molecule.

Relative to the art rejections, examiner first points to both Benco and Kim discussing the prior work in terms of metal ions in general. Thus the specific species is part of a larger genus that the lithium ion also belongs to. In the Kim reference the first full paragraph of page 2387 clearly states that they sought to synthesize chromogenic calixazacrown ethers *and investigate their complexation behavior toward alkali metal ions*. This clearly sets the scope of the reference to be the groups of alkali metal atoms of which lithium is a member. An apparent admission or statement that lithium, sodium and potassium ions are extremely similar in structure and properties as setting forth what is known or expected relative to the structure and properties of these ions. One of the differences that is known is the difference in the size of the ions (see the ionic radii from the table in the cited Sienko et al reference). This clearly points to size as a way of selectively distinguishing between the ions. These facts are relevant and important to the combination of references used in the art rejection. Examiner now points to the fact that both the Benco and Kim papers developed the calixarene azacrowns based on the structure of azacrown ethers. Both tested the complexation abilities with a series of alkali metals (see the figures and tables of both papers. Both series included lithium, sodium and potassium. Thus the tests performed were looking to see the complexation abilities of the compounds with all of the metals and found that the complexation with potassium was the strongest for compound 3 of Kim and compound II of Benco. Both Benco and Kim recognize that one reason involved in the complexation behavior is the size of the cavity. In particular applicant has pointed to the statement on page 37 of the Benco reference that clearly teaches that “selectivity is controlled primarily by a size fit effect and steric effects from the propyl substituents appended to the two rotated aryl rings of the calyx[4]arene” (see page 11 of the response). Thus Benco clearly

identifies the source of the selectivity control. The Pacey reference teaches two important things: the measurement of lithium in biological fluids is important and an azacrown having a smaller cavity than the model compounds of Benco and Kim complexes well with lithium. Thus one of skill in the art would have expected that the calixarene azacrown ethers of Kim and Benco with smaller azacrown portions consistent with the size difference between the ring size of the Kim and Benco model compounds and the Pacey compound would have selectivity towards lithium based on at least the size aspect. Also the person of skill in the art would have done it due to the desire to measure lithium as taught by Pacey and the potential for a 3-D encapsulation providing an optimized structure as taught by Kim. These same reasons apply to the obviousness-type double patenting rejection. Kim is clearly not directed to a potassium ion sensor as its sole purpose and Benco as well was not directed to a potassium ion sensor. Examiner states that if the complexation behavior of the Benco and Kim compounds were different, the papers would have reported the different results. Thus that papers are reporting the results and the focus of the papers is based on the results. As an example of this, applicant is directed to the results of compound 4 of Kim which are different from compound 3. The Kim paper presents the results of both compounds and discusses them also, even though they are different. In the same way the Benco paper is reporting the results of the tests Benco conducted. The fact that potassium is a metal that there is a desire to measure simply enhances that value of the results and allows Benco to focus the discussion in that direction. Kim on the other hand is more generally focused and shows more of the scope of understanding by those of the art. Therefore, since there is a desire to find compounds for measuring lithium as shown by Pacey, there is motivation to make the suggested change based on the difference in ring size between the azacrown ethers of Benco or Kim and Pacey. For these reasons the Barnard reference does not need to modify the compounds of Kim or Benco in that manner. It does clearly teach benefits of placing compounds intended to measure metal ions in various formats. That is how it has been used. Because the primary references are not of the scope that applicant claims, the modification of the Benco or Kim structures as proposed in the rejections above does not change the functionality of complexing with an alkali metal ion and lithium, sodium or potassium specifically as seen by the metals used to test the compounds of each reference.

Another way that examiner has responded to the arguments is to use a new ground of rejection with Pacey as the primary reference since it clearly is directed to lithium. In this rejection, the added ability to complex the target molecule as recognized, taught and/or expected by Benco and Kim provides the motivation to modify the Pacey structure to include the calixarene in addition to the azacrown of Pacey.

The other thing that needs to be pointed out is that there is no limitation relative to the selectivity of the claimed compounds for lithium over sodium or potassium. Thus the compounds claimed do not have to be selective of lithium over sodium or potassium. In this light, if the differences between the structure of the compound claimed in claim 16 and the structure that examiner believes is suggested by the combination of references -- the compound of claim 16 minus the t-butyl groups -- is the difference between a useful selectivity for lithium as disclosed in the instant specification and a compound without sufficient selectivity for lithium, there is clearly not sufficient enablement for the scope of the instant claims. If the compound that examiner feels is clearly obvious in view of the references does have sufficient selectivity of lithium over sodium and potassium, then the art rejection covers the specifically exemplified compound.

Any inquiry concerning this communication or earlier communications from the examiner should be directed to Arlen Soderquist whose telephone number is (571) 272-1265. The examiner's schedule is variable between the hours of about 6:30 AM to about 5:00 PM on Monday through Thursday and alternate Fridays.

A general phone number for the organization to which this application is assigned is (571) 272-1700. The fax phone number to file official papers for this application or proceeding is (571) 273-8300.

Information regarding the status of an application may be obtained from the Patent Application Information Retrieval (PAIR) system. Status information for published applications may be obtained from either Private PAIR or Public PAIR. Status information for unpublished applications is available through Private PAIR only. For more information about the PAIR system, see <http://pair-direct.uspto.gov>. Should you have questions on access to the Private PAIR system, contact the Electronic Business Center (EBC) at 866-217-9197 (toll-free).



Arlen Soderquist  
Primary Examiner